

<p>Temperature a solid becomes a liquid.</p> <p>Minimum number of reactants for a chemical reaction</p> <p>Copper Chloride</p> <p>Scratch test</p> <p>Feet</p> <p>Malleable</p> <p>Rust</p> <p>Inches</p> <p>You evaluate hardness with a</p> <p>Properties</p> <p>Inference</p>	<p>condense</p> <p>Is not a chemical reaction.</p> <p>In the Statue of Liberty modeling lab, the copper acetate was a</p> <p>Word Equation</p> <p>conserved</p> <p>The process of making a solution.</p> <p>Same material all the way through.</p> <p>Methane + Oxygen → Carbon Dioxide and Water</p> <p>in water.</p> <p>Closed system</p> <p>Solid part of a solution.</p> <p>piils</p> <p>Changes shape when pressure is applied and retains the shape.</p> <p>Freezing Point</p> <p>Soap is soluble _____ changes during a chemical reaction</p> <p>Rubbing alcohol</p>	<p>ml/g</p> <p>Density=</p> <p>The process of making a solution.</p> <p>g/ml</p> <p>Amount of space taken up by an object.</p> <p>Distillation</p> <p>Units for density</p> <p>Hawaiian Punch sugar and water is a solution which is a type of</p> <p>Melting Point</p> <p>Hydrogen and oxygen molecules</p> <p>Substance definition.</p>	<p>mg/g</p> <p>A method to separate a solution.</p> <p>Density of water.</p> <p>The process of making a solution.</p> <p>In a chemical reaction</p> <p>Volume ÷ Mass</p> <p>Magic</p> <p>Solid part of a solution.</p> <p>Dissolving</p> <p>To use electricity to break apart water.</p>	<p>10</p> <p>Solvent</p> <p>Density=</p> <p>The process of making a solution.</p> <p>g/ml</p> <p>Amount of space taken up by an object.</p> <p>Product of burning magnesium.</p> <p>When substances cannot enter or leave a system.</p> <p>0</p> <p>Distillation</p> <p>Units for density</p> <p>Hawaiian Punch sugar and water is a solution which is a type of</p> <p>Melting Point</p> <p>Hydrogen and oxygen molecules</p> <p>Substance definition.</p>	<p>When substances can enter or leave a system.</p> <p>Magnesium Oxide</p> <p>mixture.</p> <p>Solute</p> <p>Amount of matter an in object.</p> <p>12</p> <p>Round</p> <p>Before a chemical reaction you have 12 atoms. After the reaction you have ___ atoms.</p> <p>Open System</p> <p>Liquid part of a solution.</p> <p>Particles slowly vibrate in place.</p> <p>.996 g/ml</p> <p>Electrolysis of water</p> <p>Dissolving</p> <p>Gas</p> <p>Shape</p> <p>If you cool a hot gas it will</p>
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